**7th Grade Practice Test Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

Use the following answer choices to answer questions 1-7.

1. Physical property
2. Physical change
3. Chemical property
4. Chemical change
5. The ability of an object to catch fire is classified as a
6. Boiling point is classified as a
7. Texture is classified as a
8. Corrosion is classified as a
9. Mass is classified as a
10. Freezing is classified as a
11. Digesting is classified as a
12. Evidence of a chemical change?
13. Smoke from a marshmallow cooked over a campfire
14. Cutting your hair
15. Breaking a crayon into several pieces
16. Water freezing into ice cubes
17. Which of following activities we conducted in class is NOT a chemical change?

a. burning sugar

b. putting Alka seltzer in water

c. dissolving sugar in water

d. mixing water and calcium chloride and baking soda

1. When a log is burnt in a fire what is the evidence of the chemical change?
	1. A solid is formed.
	2. Color change of the log
	3. Temperature decreases in the area around the fire
2. All chemical reactions involve
	1. The formation of a precipitate
	2. The formation of a gas
	3. A change in color of the reactants
	4. Either absorption or release of energy
3. When a piece of paper burns, smoke and ashes are produced. Which of the following statements is true?
	1. Some of the mass is transformed into a gas
	2. Some of the mass is destroyed completely
	3. All of the mass is transformed into gas
	4. All of the mass remains in solid form
4. Iron combines with oxygen in the air to form rust. Rust is an example of a(n)
	1. Element
	2. Subscript
	3. Product
	4. Reactant
5. The substances always present at the beginning of a chemical reaction are called
	1. Catalysts
	2. Precipitates
	3. Products
	4. Reactants
6. In a chemical reaction, the mass of the products is always
	1. Less than the mass of the reactants
	2. Equal to the mass of the reactants
	3. Slightly more than the mass of the reactants
	4. Twice as much as the mass of the reactants
7. To balance a chemical reaction, you can change
	1. Compounds
	2. Subscripts
	3. Coefficients
	4. Elements
8. In the equation N2 + 3H2  2NH3, the coefficient of the hydrogen reactant is
	1. 1
	2. 2
	3. 3
	4. 6
9. An exothermic reaction is typically marked by a(n)
	1. decrease in concentration
	2. Increase in temperature
	3. Increase in elements
	4. decrease in compounds
10. Which statement reflects the law of conservation of mass?
	1. Mass can be created but not destroyed.
	2. Mass can be destroyed but not created.
	3. Mass cannot be created or destroyed.
	4. Mass can created and destroyed.
11. What is the ratio of carbon atoms to oxygen atoms in the compound with the chemical formula CO2?
	1. 1:1
	2. 2:1
	3. 1:2
	4. 2:2
12. During a chemical change you also have changes in which of the following?
	1. Physical properties
	2. Chemical properties
	3. All of the above
	4. None of the above